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Chemistry Study Guides - SparkNotes

Practice Problems: Stoichiometry. Balance the following chemical reactions: Hint a. $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$ b. $\text{KNO}_3 \rightarrow \text{KNO}_2 + \text{O}_2$ c. $\text{O}_3 \rightarrow \text{O}_2$ d. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ e. $\text{CH}_3\text{NH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{N}_2$ Hint f. $\text{Cr}(\text{OH})_3 + \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + \text{H}_2\text{O}$ Write the

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balanced chemical equations of each reaction:

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A co-founder of Alcoholics Anonymous tells how members recover and how the society functions

12.1 The Arithmetic of Equations 12

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Chapter 12 Stoichiometry127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This

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section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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Practice Problems: Stoichiometry

Chemistry. From aluminum to xenon, we explain the properties and composition of the substances that make up all matter.

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Stoichiometry 353 12.1 FOCUS Objectives 12.1.1 Explain how balanced equa-tions apply to both chemistry and everyday situations. 12.1.2 Interpret balanced chemical equations in terms of moles, representative particles, mass, ... Guided Reading and Study Workbook, Section 12.1 ...

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