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large numbers of atoms and molecules in a sample.
The number of moles of a substance has a...

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atomic mass is the mass of 1 atom of a particular element and molar mass is the mass of 1 mole of an atom or molecule how many atoms in 1 mole of water? $6.022 \times 10^{23} \times 3$

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the SI Base unit used to measure the amount of a substance. It is the number of representative particles, carbon atoms, in exactly 12 g of pure carbon 12 We use the term the Mole to represent Avogadro's number A mole of atoms is 6.022×10^{23} atoms,

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In chemistry, Avogadro's number defines the number of particles, such as atoms, ions or molecules, in one mole of a particular substance. According to the definition, Avogadro's number, N_A , is...

Bing: Chemistry Holt Study Guide Mole

The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP), one mole of any gas occupies approximately 22.4 liters.

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Amedeo Avogadro's work on gases helped define a new unit of measurement for chemistry and physics: the mole. A mole of a particular substance is equal to the number of atoms in exactly 12 g of the carbon-12 isotope. Experiments established that number to be 6.022142×10^{23} particles. Avogadro's number = N_A
 $= 6.022 \times 10^{23}$ items/mole

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